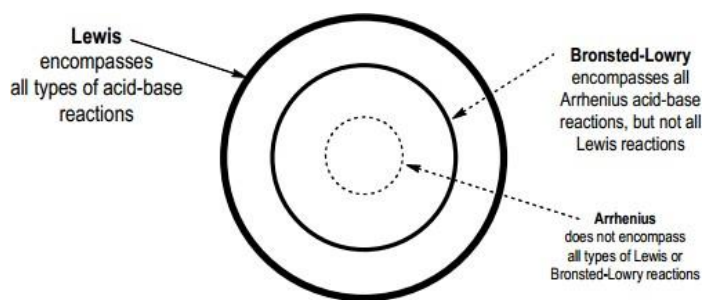


CONCEPT: ACIDS AND BASES

Before we get started: *Remember* your strong six!

(HCl, HBr, HI, HNO₃, H₂SO₄, HClO₄)



☐ Lewis definition: Same as _____

• Acid is an electron pair _____

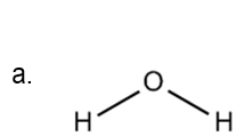
• Base is an electron pair _____

☐ Brønsted-Lowry definition:

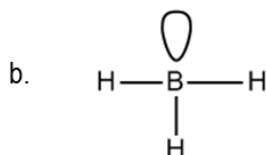
• Acid is a proton _____

• Base is a proton _____

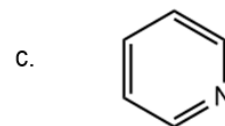
EXAMPLE: Properly identify the following molecules:



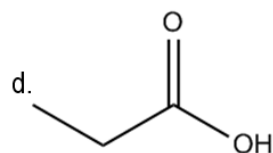
- ☐ Lewis Acid (E^+)
- ☐ Lewis Base (Nu^-)
- ☐ Brønsted Acid
- ☐ Brønsted Base



- ☐ Lewis Acid (E^+)
- ☐ Lewis Base (Nu^-)
- ☐ Brønsted Acid
- ☐ Brønsted Base



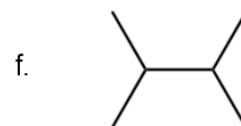
- ☐ Lewis Acid (E^+)
- ☐ Lewis Base (Nu^-)
- ☐ Brønsted Acid
- ☐ Brønsted Base



- ☐ Lewis Acid (E^+)
- ☐ Lewis Base (Nu^-)
- ☐ Brønsted Acid
- ☐ Brønsted Base



- ☐ Lewis Acid (E^+)
- ☐ Lewis Base (Nu^-)
- ☐ Brønsted Acid
- ☐ Brønsted Base



- ☐ Lewis Acid (E^+)
- ☐ Lewis Base (Nu^-)
- ☐ Brønsted Acid
- ☐ Brønsted Base

Equilibrium: A base will always attack an acid to produce _____ in the following chemical pattern:

