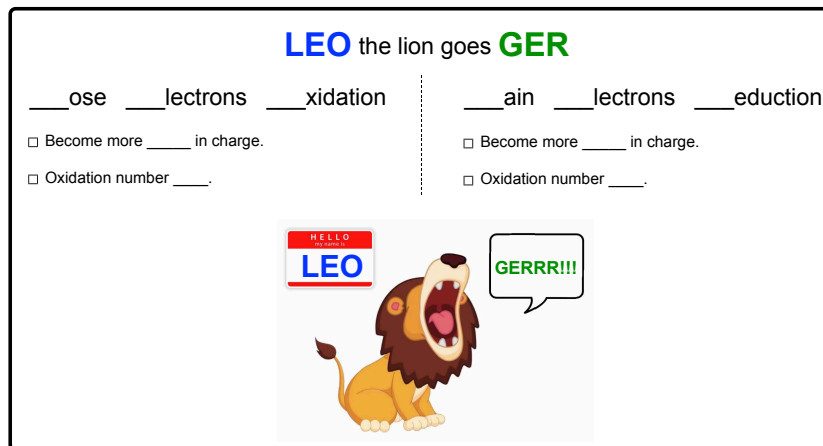


CONCEPT: REDOX REACTIONS

- **Redox Reactions** (oxidation-reduction reactions) involve transference of an electron(s) between reactants.

MEMORY TOOL



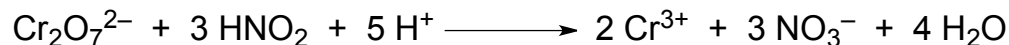
- **Reduction and Oxidation:** the _____ of oxidizing and reducing agents.
 - **Oxidizing Agent:** The element or compound that is _____.
 - **Reducing Agent:** The element or compound that is _____.

EXAMPLE: Consider the following reaction below when solid lithium reacts with the zinc ion:



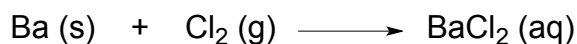
Which reactant is undergoing **oxidation** and which reactant is undergoing **reduction**?

PRACTICE: Which element is being reduced in the following reaction?

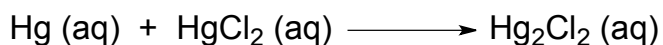


CONCEPT: REDOX REACTIONS

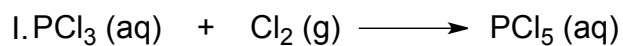
PRACTICE: Identify the oxidizing agent and reducing agent from the following redox reaction.



PRACTICE: Which element is oxidized and which is reduced in the following reaction?



PRACTICE: Which of the following represents an oxidation-reduction reaction?



a) I, II, III, and IV

b) I, II, and III

c) III

d) I and II

e) IV